

Paratene™ SHP

Stabilized Hydrogen Peroxide



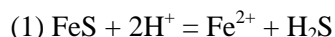
PARATENE™ SHP

“Control and Removal of Iron Sulphides and Polysulphides”

Introduction

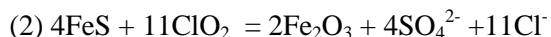
Iron sulphides and polysulphides result from the interaction of hydrogen sulphide with iron. Many Canadian oilfields have production problems directly related to the build up of iron sulphides. These problems include, plugging of water injection wells, the formation of pads in treaters, and agglomeration of deposits on contactors in amine systems.

Iron Sulphide (FeS) forms from the direct reaction of H₂S with ferrous iron (Fe²⁺) in solution. FeS is soluble in acid and will decompose into ferrous ion and H₂S in the reaction:



The dissolution of iron sulphide is further complicated when a high partial pressure of H₂S is present. Reaction (1) is reversible, and under high H₂S pressures, even strong acids will not encourage the iron sulphide particles to dissolve. This occurs rarely in surface equipment, but is a common problem in trying to remove FeS deposits formed in producing formations.

Iron Polysulphide (FeS_x) represents an entirely different problem. Polysulphides form because both iron and sulphur have multiple oxidation states, and can form combinations that can be either depleted or enriched in one of the two elements. The resulting crystals are not soluble in most acids (nitric being the exception), and an oxidizing agent is required to remove them. FeS and FeS_x will react with oxidizing agents in predictable ways. For example, consider the reaction of chlorine dioxide with iron sulphide:



The exact reaction of chlorine dioxide depends on pH, but in general the addition of an excess of oxidizing agent in the presence of iron sulphides will result in the formation of iron oxides and either elemental sulphur or sulphate ion.

Common Industry Oxidizing Agents

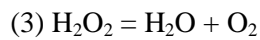
Various oxidizing agents have been considered or used to dissolve or convert iron sulphide. Several of these are listed below.

Chlorine Dioxide is a yellowish gas that can form a solution with water. Chlorine Dioxide does not act in the same manner as sodium hypochlorite, but more like hydrogen peroxide. Chlorine dioxide is toxic, and can form explosive mixtures with air. Chlorine dioxide undergoes the reactions shown in equation 2. Chlorine dioxide is also fairly corrosive.

Sodium Hypochlorite (bleach) is a powerful and common oxidizing agent. It reacts slowly with iron sulphides and forms iron oxides and elemental sulphur. Bleach is also exceptionally corrosive.

Potassium Permanganate (KMnO₄) is used in industrial and municipal water treating to remove H₂S and to treat Iron Sulphides. The reactions of KMnO₄ result in the formation of elemental sulphur or sulphate, iron oxide, and manganese dioxide. KMnO₄ is an effective method for dissolving finely dispersed iron sulphides or dissolved hydrogen sulphides, but works poorly on iron sulphide deposits, as it forms an insoluble layer of manganese dioxide on the surface of the deposit. The formation of manganese dioxide prevents penetration of the KMnO₄ to the active deposit surface, limiting the amount of the deposit that may be dissolved. The manganese dioxide formed is also an additional insoluble deposit. It can be removed by reacting it with a strong acid, but it will react with hydrochloric acid to form chlorine gas.

Hydrogen Peroxide has been used for the removal of iron sulphides and the control of hydrogen sulphide in industrial and municipal water treating. In the presence of metal ions in solution, hydrogen peroxide decomposes exothermically to form water and oxygen. This effect limits effectiveness of hydrogen peroxide in oilfield applications, as the oxygen does not react well with either the H₂S or the iron sulphides.



Paratene™ SHP incorporates a stabilizer system into hydrogen peroxide. As a result, the hydrogen peroxide does not decompose in the presence of metal ions, and more hydrogen peroxide is available to react with the hydrogen sulphide and iron sulphides. Paratene™ SHP immobilizes the iron as an acid soluble salt, and prevents any reaction with the hydrogen peroxide.

Relative Rate of Gas Evolution SHP vs. Hydrogen Peroxide

The rate of gas evolution is dependant on the temperature, the concentration of the solution and the presence of iron as iron sulphide (FeS) or polysulfide (FeS₂). The experimental results shown below give the peak rates for 175 mls of a 5% solution of hydrogen peroxide or a 5% active solution of SHP.

		Gas Evolution Rate millilitres/minute	
		SHP	H ₂ O ₂
FeS	Temperature °C		
	25	0.116	0.33
45	0.5	187	
FeS ₂	25	.09	2
	45	.09	57
	60	.16	300

At low temperatures little or no decomposition of the SHP or Hydrogen peroxide is observed. When the temperature is raised above the critical decomposition temperature (around 35°C) the rate of gas evolution increases rapidly.

Corrosion on Carbon Steel

Corrosion testing for Paratene™ SHP on carbon steel resulted in a very low corrosion rate of 23 mpy. Tests were performed at 20°C, and no decomposition of the hydrogen peroxide could be observed when Paratene™ SHP was in contact with clean carbon steel for a 24-hour period.

Relationship Between Iron and Sulphur in Peroxide Requirements

The ratio of hydrogen peroxide to iron sulphide is ½ mole of hydrogen peroxide per mole of iron, and 4 moles of hydrogen peroxide per mole of sulphur. Thus:

Sulphide	MW	H ₂ O ₂ Moles	Wt H ₂ O ₂ /g sulphide
FeS	87.85	4.5	1.74
FeS ₂	119.85	8.5	2.41
Fe _{1-x} S	<87.85	<4.5	<1.74

The amount of scale that can be dissolved is dependant on the relative ratio of iron to sulphur. For a high sulphur scale such as a pyrite, the relative amount of peroxide required goes up, for a low sulphur scale such as Troilite (FeS) or Pyrrhotite (Fe_{1-x}S) the relative amount of peroxide required goes down.

Sulphide	Grams sulphide dissolved Per litre of 5% SHP	Grams sulphide dissolved Per litre of 10% SHP
FeS	28.73	57.46
FeS ₂	20.74	41.48
Fe _{1-x} S	>28.73	>57.46

As an example, one gram of our baseline referenced iron sulphide material with 100 ml of Paratene™ SHP (5%) will yield the following results:

Chemical Formula	Abundance
Fe _(1-x) S	45-55%
FeS	25-35%
Fe ₃ O ₄	10-15%
FeS ₂	1-5%
Fe	1-5%

The initial reaction was carried out at room temperature over night. Based on the analysis, only 80 – 85% of the sample was an iron sulphide capable of reacting with SHP. The weight of the conversion products was 1.9016 g and the amount of residue after reaction with dilute HCl (cold to dissolve only the conversion products) was 0.128 or 12.8%. The conversion of the iron sulphides was essentially complete. By comparison, reaction with 5% hydrogen peroxide at 45°C for 12 hours, the weight of the sample was essentially unchanged. Some conversion of the black iron sulphide to red Fe₂O₃ was observed. Reaction with acid dissolved 75% of the sample with copious release of H₂S.

When iron pyrite was used instead of reference sulphide the results were fairly similar. When the samples were digested in SHP for 12 hours and then washed with dilute hydrochloric acid, 85% of the pyrite dissolved. When 5% hydrogen peroxide was used, 8% of the sample dissolved after treatment with hydrochloric acid. By comparison potassium permanganate dissolved 30% of the sample after the MnO₂ was removed by washing with sulphuric acid.

Discussion

The use of hydrogen peroxide is severely limited by the exothermic catalytic decomposition reaction. The rate that decomposition occurs is temperature dependant and limits the amount of hydrogen peroxide to act as an oxidizing agent against iron sulphide or other substrate. When the peroxide is stabilized against decomposition, the reactions with iron sulphides proceed to completion. This improvement of reactivity is most evident when the dissolution of iron pyrite is examined. The pyrite sample was 100% insoluble in non-oxidizing acids such as hydrochloric or sulphuric acid. When treated with stabilized peroxides 85% of the sample dissolved. Similar systems using permanganate failed to dissolve the sample and left behind a substantial amount of insoluble manganese dioxide.

Applications of Paratene™ SHP

Oil and Gas Well Cleanouts

Paratene™ SHP can be applied as a batch treatment in oil and gas wells to dissolve polymer based drilling mud. The treatment volume of Paratene™ SHP depends on the extent of the problem, the geometry of the injection zone, and the porosity of the zone. Paratene™ SHP is blended as a finished product and is applied undiluted. The treatment volume of Paratene™ SHP should be pumped into the well followed by an equal volume of 15% hydrochloric acid

Water Injection Wells

Paratene™ SHP can be applied as a batch treatment in water injection wells. The treatment volume of Paratene™ SHP depends on the extent of the problem, the height of the injection zone, and the porosity of the zone. Paratene™ SHP can be applied either undiluted, or diluted at 1:1 with fresh water. The treatment volume of Paratene™ SHP should be pumped into the well followed by a slug of water, followed by an equal volume of 15% hydrochloric acid.

Treaters and Flow Lines

Paratene™ SHP can be continuously applied at low concentrations to prevent the formation and accumulation of iron sulphides from forming in treaters and flow lines. The amount of Paratene™ SHP to add is determined by the concentration of H₂S and iron sulphide. The following chart can be used as a guideline to determine the recommended dosages.

Industrial Cleaning Applications

Paratene™ SHP can be applied to remove iron sulphides, iron polysulphides, polythionic acids and pyrophoric iron sulphides in industrial systems such as amine contactors, distillation towers, glycol systems, municipal water treating systems, and other systems where hydrogen sulphide may be present. To apply Paratene™ SHP the system should be filled with a solution of water and 30 – 50% of Paratene™ SHP. The solution should be heated to a temperature of 35 – 40°C and circulated for 4 – 5 hours. The amount of Paratene™ SHP required will depend on the total amount of expected deposit, and can be calculated from the graphs given for oilfield application. Additional Paratene™ SHP can be added to the system as the job progresses. Concentrations of more than 5% hydrogen peroxide must be avoided, as higher concentrations of H₂O₂ can be corrosive to carbon steel

***Paratene™ SHP** revolutionizes the application of Hydrogen Peroxide in the removal of iron sulfide and iron polysulfide. It provides a new and exceptionally controlled way to deliver the strong oxidizing power of peroxides in oilfield and industrial applications.*

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